

CHUKA



UNIVERSITY

UNIVERSITY EXAMINATIONS

RESIT/SPECIAL EXAM

EXAMINATION FOR THE AWARD OF DEGREE OF BACHELOR  
OF SCIENCE IN BIOCHEMISTRY

**BIOC 201: PHYSICAL BIOCHEMISTRY**

**STREAM: BSC (BIOC) Y2S1**

**TIME: 2 HOURS**

**DAY/DATE: TUESDAY 16/11/2020**

**2.30 P.M. – 4.30 P.M.**

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**INSTRUCTIONS**

- (i) Answer Question ONE and any TWO questions
- (ii) Do not write on the question paper

**QUESTION ONE (30 MARKS)**

**Constants**

$\Delta G^{\circ}$  for fructose-1-phosphate hydrolysis =  $-16\text{KJmol}^{-1}$

$R=8.315 \times 10^{-3}\text{KJmol}^{-1}\text{K}$

$T=25^{\circ}\text{C}$

**QUESTION ONE (30 MARKS)**

- a. In human erythrocytes the concentration of ATP, ADP and Pi are 2.25, 0.25 and 1.65 mM respectively, calculate the actual free energy of hydrolysis ( $\Delta G_p$ ) of ATP in the erythrocyte cell at standard pH and temperature (5 marks)
- b. Explain the fates of the actual free energy change ( $\Delta G$ ) with regard to chemical reaction (5 marks)
- c. Explain why  $\Delta G$  for favorable processes is always a negative value. (5 marks)

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- d. Explain the relevance of the first and second laws of thermodynamics to biological systems. (5 marks)
- e. ATP is usually hydrolysed in cells according to the equation  $\text{ATP} \rightleftharpoons \text{ADP} + \text{P}_i$ . Given that  $[\text{ATP}] = 1 \times 10^{-7} \text{M}$ ,  $[\text{ADP}] = 1.65 \times 10^{-1} \text{M}$  and  $[\text{P}_i] = 1 \times 10^{-1} \text{M}$ , calculate:
- The equilibrium constant. (5 marks)
  - $\Delta G^0$  for ATP hydrolysis. (5 marks)

### QUESTION TWO (20 MARKS)

- a) Using an example, describe how ATP energizes active transport (10 marks)
- b) Describe the nucleophilic displacement reactions of ATP. (10 marks)

### QUESTION THREE (20 MARKS)

- a) Explain the role of myokinase in the production of AMP during muscle contraction. (10 marks)
- b) An enzymatic hydrolysis of Fructose-1-Phosphate,  $\text{Fructose-1-PO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{Fructose} + \text{P}_i$ , was allowed to proceed to equilibrium at 25°C. The original concentration of Fructose-1-Phosphate was 0.2M, but when the system had reached equilibrium the concentration of Fructose-1-Phosphate was only  $6.52 \times 10^{-5} \text{M}$ . Calculate the equilibrium constant for this reaction and the free energy of hydrolysis of Fructose-1-Phosphate. (10marks)

### QUESTION FOUR (20 MARKS)

- a) Explain the biochemical basis for glowing of the firefly (10 marks)
- b) Outline and explain the Ping – Pong mechanism (double - displacement) of nucleoside diphosphate kinase (10 marks)
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